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section 20 1 the meaning of oxidation and reduction pages

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this section explains oxidation and reduction in terms of the loss or gain of electrons and describes the characteristics of a redox reaction it also explains how to identify oxidizing and reducing agents what are oxidation and reduction pages 631 635 1

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for each of the following balanced oxidation reduction reactions i identify the oxidation numbers for all the elements in the reactants and products and ii state the total number of electrons transferred in each reaction a 2 $\text{mno}_4^- + 1\text{aq}^- + 3\text{s}^{2-} + 1\text{aq}^- + 4\text{h}_2\text{o} + 12\text{h}^+ + 3\text{s} + 2\text{mno}^{2+} + 8\text{oh}^- + 1\text{aq}^- + 433$

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an oxidation reduction reaction is any chemical reaction in which the oxidation number of a molecule atom or ion changes by gaining or losing an electron redox reactions are common and vital to some of the basic functions of life including photosynthesis respiration combustion and corrosion or rusting

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add 3 ml of 0.1M copper(II) nitrate to each test tube in one tube place a strip of copper in another place a strip

of lead in the third place a strip of zinc put the tubes in a test tube rack repeat step 2 labeling the tubes 0 1m pb no3 2 and using 0 1m lead ii nitrate in place of copper ii nitrate

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oxidation reduction reactions redox reactions are comprised of two parts a reduced half and an oxidized half that always occur together the reduced half gains electrons and the oxidation number decreases while the oxidized half loses electrons and the oxidation number increases

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just remember that oxidation is any reaction that s gonna involve the increase in the oxygen content off a molecule now that doesn t mean that you re actually adding oxygen s it could just mean that you have more bonds two and oxygen

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increase in the oxidation number of an element section 20 3 balancing redox equations pages 645 654 this section explains how to use the oxidation number change and half reaction methods to balance redox equations pearson education inc publishing as pearson prentice hall

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carefully clean and dry each piece of leftover metal and place it in the appropriate recycling container clean the
small scale reaction surface by absorbing the contents onto a paper towel rinse it with a damp paper towel and dry
it dispose of the paper towels in the waste bin

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during a reaction between a metal and a nonmetal electrons are transferred from atoms of the to atoms of the
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zn is oxidized oxidation number 0 2 h is reduced oxidation number 1 0 al is oxidized oxidation number 0 3 cu 2 is reduced 2 0 this is not a redox reaction because each element has the same oxidation number in both reactants and products o 2 h 1 c 4

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what is the oxidation state of the indicated atoms in the following structure 12 practice problem determine using oxidation numbers if the given reaction is a redox reaction 13 practice problem assign oxidation numbers and identify if the given chemical reaction is redox or not 14 practice problem

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