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thermodynamics is a branch of physics that deals with heat work and temperature and their relation to energy entropy and the physical properties of matter and radiation

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thermodynamics science of the relationship between heat work temperature and energy thermodynamics deals with the transfer of energy from one place to another and from one form to another the key concept is that heat is a form of energy corresponding to a definite amount of mechanical work

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laws of thermodynamics four relations underlying thermodynamics the branch of physics concerning heat work temperature and energy and the transfer of such energy the first and second laws were formally stated in works by german physicist rudolf clausius and scottish physicist william thomson

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the laws of thermodynamics are a set of scientific laws which define a group of physical quantities such as temperature energy and entropy that characterize thermodynamic systems in thermodynamic equilibrium the laws also use various parameters for thermodynamic processes such as thermodynamic work and heat and establish relationships

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thermodynamics is the study of the relationship between heat or energy and work in other words thermodynamics looks at how we can put energy into a system whether it is a machine or a molecule and make it do work

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thermodynamics is the study of energy and heat the laws of thermodynamics describe the relationship between matter and energy and how they relate to temperature and entropy many texts list the three laws of thermodynamics but really there are four laws although the 4th law is called the zeroeth law

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thermodynamics is the field of physics that deals with the relationship between heat and other properties such as pressure density temperature etc in a substance specifically thermodynamics focuses largely on how a heat transfer is related to various energy changes within a physical system undergoing a thermodynamic process

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the first law of thermodynamics applies the conservation of energy principle to systems where heat transfer and doing work are the methods of transferring energy into and out of the system the first law of thermodynamics states that the change in internal energy of a system Δ u equals the net heat transfer into the system q plus the

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define the first law of thermodynamics describe how conservation of energy relates to the first law of thermodynamics identify instances of the first law of thermodynamics working in everyday situations including biological metabolism

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there are three types of systems in thermodynamics open closed and isolated an open system can exchange both energy and matter with its surroundings the stovetop example would be an open system because heat and water vapor can be lost to the air a closed system on the other hand can exchange only energy with its surroundings not matter

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thermodynamics study of the relationships among heat work temperature and energy any physical system will spontaneously approach an equilibrium that can be described by specifying its properties such as pressure temperature or chemical composition if external constraints are allowed to change these properties generally change

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the determination of the relationships among the various properties of materials without knowing their internal structure is the subject of thermodynamics historically thermodynamics was developed before an understanding of the internal structure of matter was achieved

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thermodynamics in physics is a branch that deals with heat

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work and temperature and their relation to energy radiation and physical properties of matter to be specific it explains how thermal energy is converted to or from other forms of energy and how matter is affected by this process thermal energy is the energy that comes from heat

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law of thermodynamics the first law of thermodynamics states that energy can be converted from one form to another with the interaction of heat work and internal energy but it cannot be created nor destroyed under any circumstances mathematically this is represented as $\Delta u \ q \ w$ 1 1 $\Delta u \ q \ w$ with

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thermodynamics heat energy laws in order to carry through a program of finding the changes in the various thermodynamic functions that accompany reactions such as entropy enthalpy and free energy it is often useful to know these quantities separately for each of the materials entering into the reaction

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updated on july 28 2019 the laws of thermodynamics are important unifying principles of biology these principles govern the chemical processes metabolism in all biological organisms the first law of thermodynamics also known as the law of conservation of energy states that energy can neither be created nor destroyed

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