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in chemistry we define equilibrium as the point at which the rate of the forward reaction is equal to the rate of the reverse reaction we learned about rates in unit five as how quickly a reaction proceeds

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decreases in response to a change an equilibrium shifts that is the equilibrium favors either the forward or the reverse reaction when a change is introduced to a system in equilibrium the equilibrium shifts in the direction that relieves the change this rule was first observed by henri le châtelier shown in figure 26 today the

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various methods can be used to solve the two fundamental types of equilibrium problems 1 those in which we calculate the concentrations of reactants and products at equilibrium and 2 those in which we use the equilibrium constant and the initial concentrations of reactants to determine the composition of the equilibrium mixture

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there are two fundamental kinds of equilibrium problems those in which we are given the concentrations of the reactants and the products at equilibrium or more often information that allows us to calculate these concentrations and we are asked to calculate the equilibrium constant for the reaction and

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in this chapter you will learn how to predict the position of the balance and the yield of a product of a reaction under specific conditions how to change a reaction s conditions to increase or reduce yield and how to evaluate an equilibrium system s reaction to disturbances 5 1 prelude to equilibrium

calculating equilibrium constant K_p using partial pressures

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google classroom definition of equilibrium constant K_p for gas phase reactions and how to calculate K_p from K_c key points the equilibrium constant K_p describes the ratio of product and reactant concentrations at equilibrium in terms of partial pressures for a gas phase reaction $a_a g + b_b g \rightleftharpoons c_c g + d_d g$

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what is true about a system in chemical equilibrium the amount of reactant changes but over time remains about the same for the reaction $H_2CO_3 \rightleftharpoons CO_2 + H_2O$ the products at completion are present at 99 and the reactant at 1

section 7 5 quantitative changes in equilibrium systems

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1 a given H_2 0.14 mol l H_2 0.040 mol l I_2 0.010 mol l K_c 0.020 required Q direction of reaction to reach equilibrium H_2 solution Q 2 g l I_2 2 g l H_2 0.040 mol l I_2 0.010 mol l Q 0.020 statement the value of Q is 0.020 Q is equal to K_c so the reaction is at equilibrium

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